

Chemistry Terminology II (Speedy Study Guides)

Speedy Study Guides		CHEMISTRY TERMINOLOGY II	
effusion	Gas molecules in a container escape from tiny pinholes into a vacuum with the same average velocity they have inside the container. They also move in straight-line trajectories through the pinhole.		
electric charge	A property used to explain attractions and repulsions between certain objects. Two types of charge are possible: negative and positive. Objects with different charge attract, objects with the same charge repel each other.		
electric dipole	An object whose centers of positive and negative charge do not coincide. For example, a hydrogen chloride (HCl) molecule is an electric dipole because bonding electrons are on average closer to the chlorine atom than the hydrogen, producing a partial positive charge on the H end and a partial negative charge on the Cl end.		
electric field	A field of forces that act on any electric charge placed within it. The stronger the field, the stronger the force that acts on the charge. For example, the positive charge on an atomic nucleus creates an electric field that traps electrons.		
electrode	An electrically conducting surface that allows electrons to be transferred between reactants in an electrochemical cell.		
electrolytic cell	A device that uses electricity from an external source to drive a redox reaction.		
electrolysis	The process of driving a redox reaction in the reverse direction by passage of an electric current through the reaction medium.		
electrolyte	A substance that dissociates fully or partially into ions when dissolved in a solvent, producing a solution that conducts electricity. (See strong electrolyte, weak electrolyte.)		
electronegativity	Electronegativity is a measure of the attraction an atom has for bonding electrons. Bonds between atoms with different electronegativities are polar, with the bonding electrons spending more time on average around the atom with higher electronegativity.		
electrorefining	Electrorefining is a method for purifying a metal using electrolysis. An electric current is passed between a sample of the impure metal and a cathode when both are immersed in a solution that contains cations of the metal. Metal is stripped off the impure sample and deposited in pure form on the cathode.		
emission spectrum	A plot of relative intensity of emitted radiation as a function of wavelength or frequency.		
emollient	A substance added to a formulation that gives it softening ability. For example, oils that can soften skin are added as emollients in some skin creams.		
emulsion	A colloid formed from tiny liquid droplets suspended in another, immiscible liquid. Milk is an example of an emulsion.		
enantiomer	Two molecules that are nonsuperimposable mirror images of each other. One enantiomer rotates plane-polarized light to the left, the other rotates it to the right.		
endothermic	A process that absorbs heat. The enthalpy change for an endothermic process has a positive sign.		
endpoint	The experimental estimate of the equivalence point in a titration.		
energy	Energy is an abstract property associated with the capacity to do work.		
enthalpy	Enthalpy (H) is defined so that changes in enthalpy (ΔH) are equal to the heat absorbed or released by a process running at constant pressure. While changes in enthalpy can be measured using calorimetry, absolute values of enthalpy usually cannot be determined. Enthalpy is formally defined as $H = U + PV$, where U is the internal energy, P is the pressure, and V is the volume.		
entropy	Entropy is a measure of energy dispersal. Any spontaneous change disperses energy and increases entropy overall. For example, when water evaporates, the internal energy of the water is dispersed with the water vapor produced, corresponding to an increase in entropy.		
enzyme	Protein or protein-based molecules that speed up chemical reactions occurring in living things.		

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